Relative Atomic Mass

\*Mass no. of copper – 63, 65.

\*Abundance of copper – 31%, 69%

(63x69)+(65+31) = 63.62

 100

Compounds are found in the periodic table

False

Compounds are easy to separate

False

The little number tells us how much of each element

True

Relative Atomic Mass is in grams

False

Chromium has 4 energy levels:

1S

2S – 2P

3S – 3P – 3D

4s

Being half filled is more stable than having nothing in it.

Atomic orbital is the region of space around a nucleus that has a high chance of finding an electron.

2 is the maximum number of electrons that can occupy any atomic orbital.

There are 2 main energy levels and 5 orbitals in nitrogen.

The difference between 2s and 2p is the shape of the atomic orbital.

The element that is 13 on the periodic table is aliminium.

Transition metal is defined by it’s last electron being in a D orbital. (EX, chromium.)

Chlorine gains one electron and becomes a chloride ion with a negative charge.

Oxygen gains two electrons and becomes an oxygen ion with a negative charge.

Sulphur atom gains three electrons and becomes sulphide ion with a negative charge.

Sodium atom loses one electron and becomes a sodium ion with a positive charge.

Potassium atom loses one electron to become potassium ion with a positive charge.

Beryllium Atom loses two electrons and becomes beryllium ion with a positive charge.

Electrons will always be transferred from the metal to the non-metal. Every atom is looking to gain a full outer shell.

Checklist:

\*Put dots on the metal, crosses on the non-metal.

\*Arrows to show the electron moving.

\*Ions drawn out below.

\*Square brackets and charges.

\*Name of the ionic compound formed is written.